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*5 - 1 CHAPTER 5:
STRUCTURE OF
POLYMERS "The
time has come," the*

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*Walrus said, "To talk
of many things: Of
shoes--and
ships--and sealing-
wax-- Of
cabbages--and
kings--" Lewis
Carroll, Through the
Looking Glass (1872)
shoes, ship, sealing
wax, cabbage, and a
king The "many
things" listed by the
Walrus are actually
very similar in*

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chemical

composition and

structure. If we

consider only

wooden ...

[Mass of Proton -](#)

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In chemical bonds,

atoms can either

transfer or share

their valence

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electrons. In the extreme case where one or more atoms lose electrons and other atoms gain them in order to produce a noble gas electron configuration, the bond is called an ionic bond. Typical of ionic bonds are those in the alkali halides such as sodium chloride,

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NaCl.

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*Consider the
18-gauge wire. The
density of free
electrons is $8.5 \times$
 10^{28} electrons
per cubic meter. (a)
How many atoms are
in 1.00 m^3 of
copper? (b) How*

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*many free electrons
are there per copper
ato*

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Covalent Answers

Bonding Covalent

Bonding

Pairs of electrons located between atoms represent covalent bonds. Multiple pairs represent multiple bonds, such as double bonds and triple bonds. An

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alternative form of representation, not shown here, has

bond-forming

electron pairs

represented as solid

lines. [citation

needed] Lewis

proposed that an

atom forms enough

covalent bonds to

form a full (or

closed) outer

electron shell. In the

...

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The protons, neutrons, and electrons in the atoms of iron and gold are identical—there are just different numbers of them. In theory, you could turn iron into gold by taking iron atoms

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and adding 53

protons, 88

neutrons, and 53

electrons to each

one. But if that were

as easy as it sounds,

you can bet all the

world's chemists

would be very rich

indeed! But let's

suppose you could

turn atoms ...

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[Ionic Bonds -](#)

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Silicon atoms have 14 electrons, where the four electrons in the last layer are called valence electrons. In a crystal solid, each silicon atom normally shares one of its four valence

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with another silicon atom. The silicon crystal molecule is formed of 5 silicon atoms in a covalent junction. The process of doping introduces an ...

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Bonding electrons, which are shared by a pair of atoms and nonbonding electrons, which belong to a particular atom but do not participate in bonding. In CCl_4 the central carbon atom has four bonding groups of electrons. Each chlorine atom has three nonbonding

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*pairs of electrons.
The arrangement of
the atoms is correct
in my structure. That
is the carbon is the
central atom and the
four ...*

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and Aromaticity -
Vanderbilt
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16. HCN , $1 + 4 + 5$
 $= 10$ valence

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N is hybridized, both

C and N atoms are

sp hybridized. The

C-H σ bond is

formed from overlap

of a carbon sp hybrid

orbital with a

hydrogen 1s atomic

orbital. The triple

bond is composed of

one σ bond and two

π bonds. The sigma

bond is formed from

head-to-head

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5. Atoms combine in the ratio of small whole number to form compounds. 6. The relative number and kinds of atoms are constant in a given compound. 9 • Atom. Atoms are the smallest particles of

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*an element which
can take reaction.*

Size of an atom:

*Atomic radius is
measured in
nanometres. Atomic
radii of hydrogen
atom = 1×10^{-10}
m. Symbols of
atoms: (a) Symbols
for some elements as
...*

[21.6 Biological
Effects of Radiation](#)

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Carbon atoms contain six positively charged protons and six neutral neutrons within the nucleus. Six negatively charged electrons surround the nucleus. The outer electron shell in carbon has the potential to store four additional

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electrons. This crucial detail enables carbon to bond to up to four other atoms and form complex organic compounds. Lauric acid is an example of a simple lipid, which ...

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Those atoms that gain or lose electrons become charged ions that pass into the electrolyte. Those ions that gain or lose electrons to become uncharged atoms separate from the electrolyte. The formation of uncharged atoms from ions is called discharging. The

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energy required to cause the ions to migrate to the electrodes, and the energy to cause the change in ionic state, is provided by the ...

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